END OF TERM 1 EXAMS

CHEMISTRY

FORM FOUR

PAPER 1
TIME: 2 HOURS

NAME……ADM NO:……

SIGN…… INDEX NO:……

INSTRUCTIONS TO CANDIDATES

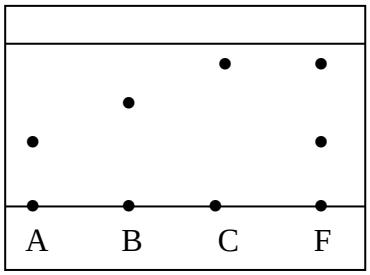
- Write your **name** and **index number** in the spaces provided above
- **Sign** and write the **date of examination** in the spaces provided.
- Answer ***all*** the questions in the spaces provided.
- Mathematical table and silent electronic calculators may be used.
- **All** working must be clearly shown where necessary.

FOR EXAMINERS USE ONLY

QUESTION	MAXIMUM SCORE	CANDIDATE'S SCORE
1-26	80	

This paper consists of 13 printed pages.
Candidates should check the question paper to ascertain all the pages are printed as indicated And no questions are missing.

1. Three pure pigments were prepared and their spots placed on a filter paper as shown below. The three pigments are A, B and C. A mixture F was also placed on the filter paper at the same time with the pure pigments. The filter paper was then dipped in ethanol solvent and left for some half an hour. The results were obtained as follows.



(i) Which of the three pure pigments is most sticky? Give a reason for your answer. (1mk)

(ii) Which pure pigment is not present in the mixture **F**? (1mk)

(iii) Show on the diagram the baseline. (1mk)

2. Describe how a pure sample of lead (II) carbonate can be prepared in the laboratory starting with lead II oxide.

(3mks)

3. Write ionic equations for the reactions between : (4mks)

(a) Aqueous solution of sodium chloride and lead nitrate

(b) Aqueous solution of barium chloride and magnesium sulphate

(c) Aqueous solution of potassium hydroxide and dilute nitric acid

(d) Zinc and an aqueous solution of copper (II) sulphate

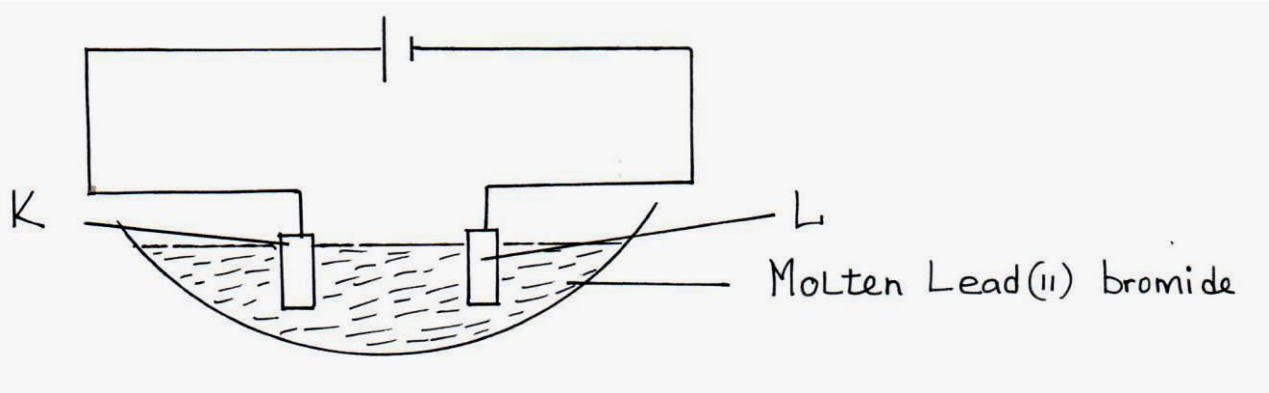
4. If it takes 20 seconds for $200cm^3$ of oxygen gas to diffuse across a porous plug. How long will it take an equal volume of sulphur (IV) oxide to diffuse across the same plug? (3mks)

5. Explain reaction of lithium, sodium and potassium with water and write down the chemical equations in each case. (6mks)

6.A mixture contains ammonium chloride, aluminium oxide and sodium chloride. Describe how each solid substance can be obtained from the mixture. (3mks)

7. State the difference between the following salts;
Deliquescent and hygroscopic salts. (2mks)

8.. Below is a set-up of apparatus used to investigate the effect of electric current on molten lead (II) bromide.



(a) Name electrode. (1mk)

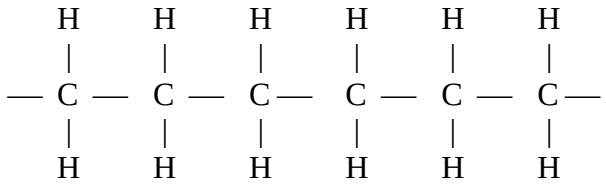
K

L

(b) State the observation made at electrode **K**. (1mk)

(c) Write an equation for the reaction taking place at electrode **L**. (1mk)

9.A sample of a polyethene polymer has the following structure.



a) Draw the structural formula of the monomer that makes the above polymer

b)The polymer is found to have a molecular mass of 2268g. Determine the number of monomers in the polymer. **(H = 1, C = 12).** (2mks)

10.The isotopes hydrogen are $^{1}_{1}H$ and $^{2}_{1}H$. Determine the molecular masses of the molecules formed when each of these isotopes react with chlorine. (Cl = 37, H=1) (1mk)

11. The table below gives the atomic numbers of elements W,X,Y and Z. The letters do not represent the actual symbol of the elements

Element	**A**	**B**	**C**	**D**
Atomic number	**9**	**10**	**11**	**12**

a) Which **one** of the elements is unreactive? Explain (1mk)

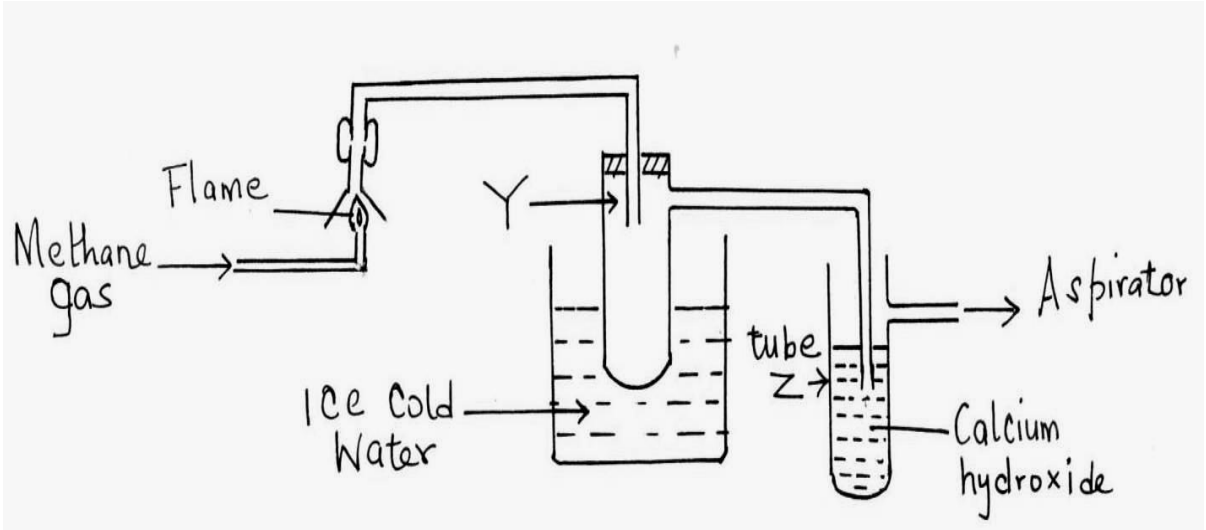
b)i) Which **two** elements would react most vigorously with each other? (1mk)

ii) Give the formula of the compound formed when the elements in b (i) above react (1mk)

12a) Distinguish between a hydrogen bond and covalent bond (2mks)

b) Explain why the boiling point of water is higher than that of hydrogen Sulphide
(Relative molecular mass of water is 18 while that hydrogen sulphide is 34) (2mks)

13. The set-up below was used to investigate the products of burning methane gas. Study it and answer the questions that follow:



(a) What product will be formed in the test tube **Y**? (1mk)

(b) State and explain the observations made in tube **Z**. (2mks)

14. Below are P^H values of some solutions.

Solution	Z	Y	X	W
P^H	6.5	13.5	2.2	7.2

(i) Which solution is likely to be

I Acidic rain. (½mk)

II Potassium hydroxide (½mk)

(ii) A basic substance **V** reacted with both solutions **Y** and **X**. What is the nature of **V**. (1mk)

15. In cold countries, salt is sprayed on the road to melt ice but in the long run it costs the motorists.
(a) How does the salt help in melting ice? (1mk)

(b) How does the salt affect the motorists? (1mk)

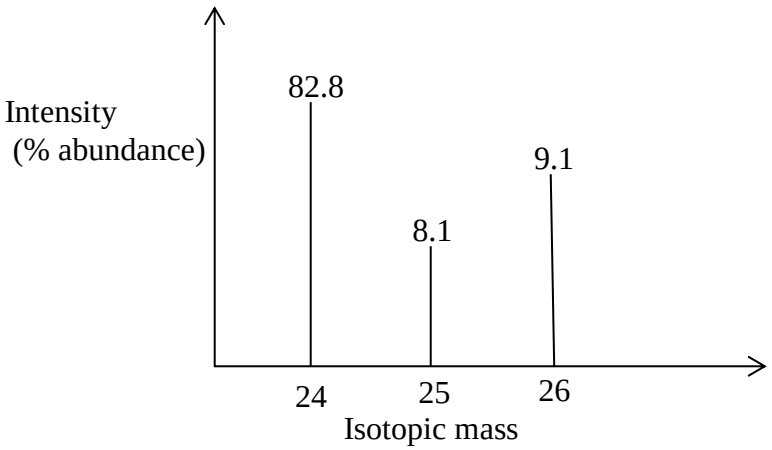
16. Using dots **(.**) and crosses (**x**) to represent electrons, show bonding in the compounds formed when the following elements react: (**Si=14, Na=11, Cl=17**).
(a) Sodium and chlorine. (2 Mks)

(b) Silicon and chlorine. (2 Mks)

17. (a) State Graham's law of diffusion. (1mk)

(a) $20cm^3$ of an unknown gas Q takes 12.6 seconds to pass through a small orifice, $10cm^3$ of oxygen gas takes 11.2 seconds to diffuse through the same orifice under the same conditions of temperatures and pressure. Calculate the molecular mass of unknown gas Q (O = 16). (3mks)

18. The peaks below show the mass spectrum of element X.



Calculate the relative atomic mass of X. (2mks)

19. Name the following compounds using the IUPAC rules.

(a) $CH_3CH_2CHCH_2CH_2CH_3$
$\quad\quad\quad |$
$\quad\quad\quad CH_2CH_3$ ____ (1mk)

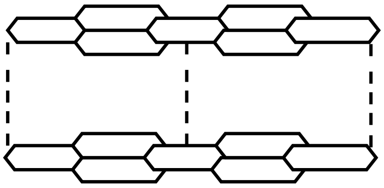
(b) $CH_3CHCHCH_3$ ____ (1mk)

(c) Draw TWO structural formulae of isomers of compound with the molecular formula **$CH_3CH_2CH_2CH_3$**

(2mks)

20.(a) What is meant by allotropy? (1 mk)

b) The diagram below shows the structure of one allotropes of carbon.



i) Identify the allotrope (1 mk)

ii) State **one** property of the above allotrope and explain how it is related to its structure. (2mk) .

21. $24cm^3$ of a solution of 0.1 M potassium hydroxide were exactly neutralized by $30cm^3$ of a solution of sulphuric acid. Find the molarity of the acid.

(3 mks)

22. (a) Give **one** use of hygroscopic substances in the laboratory. (1 mk)

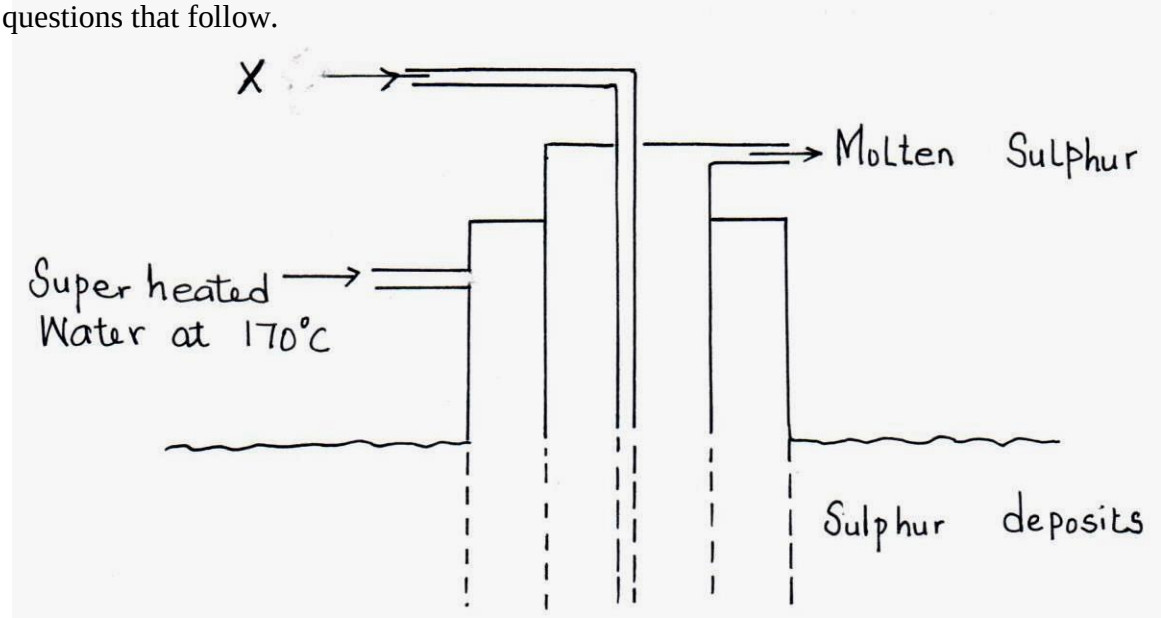
(b) What is meant by the terms: (2 mks)

(i) Isotopes

(ii) Mass number

(c) The formulae for a chloride of phosphorus is PCl_3. What is the formula of its sulphide? (1 mk)

23. The diagram below shows the Frasch process used for extraction of sulphur. Use it to answer the questions that follow.



(i) Identify **X**. (1mk)

(ii) Why is it necessary to use super heated water in this process? (1mk)

(iii) State **two** physical properties of sulphur that makes it possible for it to be extracted by this method. (1mk)

24. A certain carbonate XCO_3, reacts with dilute hydrochloric acid according to the equation given below:

$XCO_{3(s)} + 2HCl_{(aq)} \longrightarrow XCl_{2(aq)} + CO_{2(g)} + H_2O_{(l)}$

If 4g of the carbonate reacts completely with $40cm^3$ of 2M hydrochloric acid, calculate the relative atomic mass of X**. (C=12.0 ,O=16.0, Cl=35.5).** (3 Mks)

25. The table below gives some properties of three substances **I**, **J** and **K**. Study it and answer the questions that follow.

Substance	Mpt (°C)	Solubility in water	Electrical conductivity	
			Solid	Molten
I	1063	Insoluble	Conduct	Conduct
J	113	Insoluble	Doesn't	Doesn't
K	402	Sparingly soluble	Doesn't	Conduct and is decomposed

(a) Suggest the type of structure in

(i) **I** (1mk)

(ii) **K** (1mk)

Explain why the molten **K** is decomposed by electric current but **I** is not decomposed.(2mks)